

CLASS X- CHEMISTRY
METALS AND NON-METALS

ASSIGNMENT-2

MULTIPLE CHOICE QUESTION - 2.1

1. Octet rule was given by -
(A) Rutherford (B) Soddy (C) Lewis & Kossel (D) None of these
2. Exception of octet rule is -
(A) K (B) Ca (C) N (D) He
3. Ionic bond is formed by -
(A) loss of electrons only. (B) gain of electrons only.
(C) loss and gain of electrons both. (D) sharing of electrons.
4. Ionic bond is formed between -
(A) two electropositive elements.
(B) two electronegative elements.
(C) Electropositive & electronegative elements.
(D) None of these
5. During formation of ionic bond -
(A) there is force of repulsion between two negative ions.
(B) there is force of repulsion between two positive ions.
(C) there is force of attraction between positive & negative ions.
(D) none of these.
6. In the formation of ionic bond, cation is formed by-
(A) gain of electron (s). (B) loss of electron(s).
(C) sharing of electron(s). (D) None of these
7. Ionic compound have -
(A) low melting and high boiling points. (B) high melting and low boiling points.
(C) low melting and low boiling points. (D) high melting and high boiling points.
8. Ionic compounds conduct electricity in-
(A) solid state (B) fused state.
(C) gaseous state. (D) Do not conduct electricity at all.
9. Ionic compounds are soluble in-
(A) water (B) benzene (C) ether (D) alcohol
10. Physical nature of most of the ionic compounds is-
(A) solid (B) liquid (C) gas (D) May exist in any state.

SUBJECTIVE QUESTION - 2.2

1. Define octet rule.
2. Define electrovalency.
3. Explain the brittle nature of ionic compounds.
4. Why ionic compounds have high melting and boiling points ?
5. Why ionic compounds show electrical conductivity in fused or soluble state ?