CLASS X- CHEMISTRY

# CHEMICAL REACTIONS & EQUATION

## ASSIGNMENT-3

## MULTIPLE CHOICE QUESTION-3.2

1. In the reaction $Mg + Cl_2 \rightarrow MgCl_2$

   Chlorine may be regarded as -

   (A) an oxidising agent (B) a reducing agent

   (C) a catalyst (D) providing an inert medium

2. When the gases sulphur dioxide and hydrogen sulphide react, the reaction is

   $SO_2 + 2H_2S \rightarrow 2H_2O + 3S$

   Here hydrogen sulphide is acting as -

   (A) an oxidising agent (B) a reducing agent

   (C) a dehydrating agent (D) a catalyst

3. Which of the following statements is/are false for oxidation reaction?

   (A) Gain or addition of electronegative radical

   (B) Removal of hydrogen atom.

   (C) Removal or loss of electropositive radical or element

   (D) None of these

4. $CiO + H_2 \rightarrow H_2O + Cu$, reaction is an example of -

   (A) redox reaction (B) synthesis reaction

   (B) neutralisation (D) analysis reaction

5. Which of the following is an example of oxidation reaction ?

   (A) $Sn^{+2} - 2e^- \rightarrow Sn^{+4}$ (B) $Fe^{+3} + e^- \rightarrow Fe^{+2}$

   (C) $Cl_2 + 2e^- \rightarrow 2Cl$ (D) None of these

6. In the process of burring of magnesium in air, magnesium undergoes -

   (A) reduction (B) sublimation (C) oxidation (D) all of these

7. A substance which oxidises itself and reduces other is known as-

   (A) an oxidising agent (B) a reducing agent (C) Both of these (D) None of these

8. Oxidation is a process which involves -

   (A) addition of oxygen (B) removal of hydrogen

   (C) loss of electrons (D) All are correct

9. In the reaction $PbO + C \rightarrow Pb + CO$.

   (A) PbO is oxidized (B) C acts as oxidising agent.

(C) C acts as a reducing agent.

(D) This reaction does not represent a redox reaction.

10. A redox reaction is one in which -

(A) both the substances are reduced.

(B) both the substances are oxidised.

(C) and acid is neutralised by the base.

(D) one substance is oxidised, which the other is reduced.

## SUBJECTIVE QUESTION-3.2

1. Oxidation reaction have some harmful effects. Comment on the sentence.

2. Can oxidation occur without reduction ? Explain

3. Explain the terms oxidation and reduction with examples.

4. What is rancidity? Example with example.

5. What do you mean by corrosion ?

6. Identify the substances that are oxidized and the substances that are reduced in the following reactions -

(a) $ZnO + C \longrightarrow Zn + CO$

(b) $MnO_2 + 4HCl \longrightarrow MnCl_2 + 2H_2O + Cl_2$

(c) $2FeCl_3 + H_2S \longrightarrow 2FeCl_2 + S + 2HCl$

(d) $3Mg + N_2 \longrightarrow Mg_3N_2$